***Lesson Plans for the Week of: 1/9/17 Teacher: Hough Course: Chemistry Period: 9***

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| Elements of  a Lesson | **Monday** | **Tuesday** | **Wednesday** | **Thursday** | **Friday** |
| Objective/  Focus/  Essential  Question | This lesson done on 1/12/17 due to snow day  CH.3c,d  --differentiate between polyatomic ions and monatomic ions  --How to name non-binary ionic compounds, given the chemical formula | This lesson done on 1/13/17 due to snow day  CH.3a,c,d  --Write the chemical formula for non-binary ionic compounds, given the name of the compound | Delete this lesson  CH.3a,c,d and old  Review for test | Lesson planned for 1/9/17 done this day  CH.3a,c,d and old  Test | Lesson planned for 1/10/17 done this day  CH.3a,d  --identify properties of covalent (molecular) compounds  --Write the names of binary covalent compounds, given the formula |
| Lesson/Act.  Type of Presentation | Bellwork: Interpret m vs V graph: Is A or B less dense?  Define polyatomic ion—contrast with monatomic ion; give examples  Differentiate between binary compounds those with polyatomic ions: the latter contains more than 2 elements in the compound  List of common polyatomic ions on textbook p. R47  Use HCN and Mg(OH)2, particle diagrams to explain how the charge works with polyatomic ions  Individual:  SKILL:  Naming ionic compounds which contain polyatomic ions:  Guided practice: Practice naming problems 9.5 on p. 278 or ex: Mg(OH)2, NaNO3, NH4Cl;  Independent practice #16; emphasize that the metal gets the first name, except when the anion is NH4+ | Individual:  Bellwork: Name Fe(OH)3  Go over results  Whole group:  Review previous day’s work  SKILLS:  Model writing ionic compound (w/polyatomic ion) formulas:  Individual:  Guided practice: Practice problems 9.4 on p. 277  Independent practice #14 and 15a; emphasize that the metal gets the first name, except when the anion is NH4+ | Review for Test | Test | Whole group:  Define covalent compounds (molecular compound);  list properties: note that forces between the units (molecules) are weaker than the forces within the units; view crystal structures to show difference between ionic and covalently structures |
| Evaluation | Teacher observation and results of student practice | Teacher observation and results of student practice |  |  |  |
| Extension/  Homework | Textbook p. 278#17, p.298#76a,b 299#82a,c,h | Textbook p. 279#23  Pearson adapted worksheet: mixed practice, labelled |  |  |  |

MATERIALS:

Monday: Unit 6 ws3 #3,8 worksheet

Tuesday: textbook; Ionic compound worksheet from hand me down teacher folder; teacher-adapted resource worksheet from Pearson textbook chapter 9

Wednesday: textbook

Thursday: Teacher-made test

Friday: Teacher-made notes; U6ws4